

## Atomic Masses of the Elements (in amu or g/mole)

Element	Symbol	At. No.	Atomic Mass	Common ion	Name of ion
Aluminium	Al	13	26.98	$\text{Al}^{3+}$	
Arsenic	As	33	74.92		
Barium	Ba	56	137.33	$\text{Ba}^{2+}$	
Boron	B	5	10.81		
Bromine	$\text{Br}^-$	35	79.90	$\text{Br}^-$	Bromide
Calcium	Ca	20	40.08	$\text{Ca}^{2+}$	
Carbon	C	6	12.011	$\text{HCO}_3^-$ $\text{CO}_3^{=}$	Bicarbonate Carbonate
Chloride	$\text{Cl}^-$	17	35.45	$\text{Cl}^-$	
Copper	Cu	29	63.55	$\text{Cu}^{2+}$	
Fluorine	F	9	18.998	$\text{F}^-$	Fluoride
Hydrogen	H	1	1.008	$\text{H}^+$	
Iron	Fe	26	55.85	$\text{Fe}^{2+}$ ( $\text{Fe}^{3+}$ )	
Lithium	Li	3	6.94	$\text{Li}^+$	
Magnesium	Mg	12	24.305	$\text{Mg}^{2+}$	
Manganese	Mn	25	54.94	$\text{Mn}^{2+}$	
Nitrogen	N	7	14.007	$\text{NO}_3^-$ $\text{NH}_4^+$	Nitrate Ammonium
Oxygen	O	8	15.999	$\text{OH}^-$	Hydroxide
Potassium	K	19	39.10	$\text{K}^+$	
Silicon	Si	14	28.09		
Sodium	Na	11	22.99	$\text{Na}^+$	
Strontium	Sr	38	87.62	$\text{Sr}^{2+}$	
Sulphur	S	16	32.066	$\text{SO}_4^{=}$	Sulphate
Zinc	Zn	30	65.39	$\text{Zn}^{2+}$	

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Masses of some common complex ions and compounds

amu or g/mole

Element	Symbol	Atomic Mass
Ammonium	$\text{NH}_4^+$	18.04
Calcium carbonate	$\text{CaCO}_3$	100.09
Bicarbonate	$\text{HCO}_3^-$	61.02
Carbonate	$\text{CO}_3^{=}$	60.01
Carbon dioxide	$\text{CO}_2$	44.01
Nitrate	$\text{NO}_3^-$	62.01
Hydroxide	$\text{OH}^-$	17.01
Water	$\text{H}_2\text{O}$	18.02
Silica	$\text{SiO}_2$	60.09
Sulphate	$\text{SO}_4^{=}$	96.06